

## Calculating Ph Packet Answers Pogil Answer

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Calculating Ph Packet Answers Pogil Answer The pH of the solution will be governed by the following base hydrolysis equilibrium, which is just the reverse of the neutralization reaction:  $A^- + H_2O \rightleftharpoons HA + OH^-$   $K_b$  for  $A^-$ , which may be calculated from  $K_a$  of HA is  $K_b = K_w / K_a = 1.00 \times 10^{-14} / 1.00 \times 10^{-5} = 1.00 \times 10^{-9}$  Making the usual

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Keep in mind that.  $x$  is equal to the concentration of hydroxide ions now in the solution.  $[OH^-] = 6 \times 10^{-4} M$ . Use this value in the equation for pOH:  $pOH = -\log [OH^-] = -\log (6 \times 10^{-4}) = 3.2$ . Remember that the sum of pH and pOH is always 14.

*Calculating pH and pOH - High School Chemistry*

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